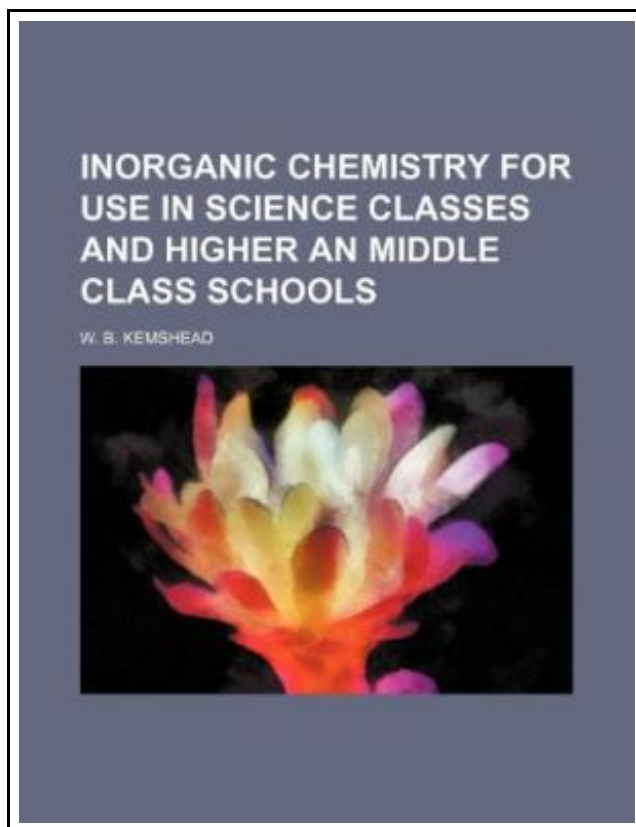


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1873 Excerpt: .weight, or about 480 times its volume of the gas, increasing in volume about one-third and acquiring a density of 1.2109. The formation of solution of hydrochloric acid is evidenced by the instantaneous change of the blue colour to red. 73. Solution of Hydrochloric Acid.--A solution of this acid in water is one of the most useful and indispensable reagents in the laboratory, and therefore its preparation, and properties, and tests for its purity, become matter of great importance to the chemist. It is most conveniently prepared by the action given in the second equation on page 92. Place one part by weight of dry sodic chloride (common salt) previously fused, and two parts by weight of sulphuric acid (previously diluted with about one third of water) in a capacious flask or retort, shake it well together, and connect it with two (or even three) Woulfe's bottles, and apply a gentle heat. In the first Woulfe's bottle is placed a little water, to arrest the sulphuric acid which may be carried over, as well as any other impurities which are sometimes mechanically carried over: with the gas. The second bottle may contain water to dissolve the gas, and form the solution. The following precautions are, however, necessary. The bottle containing the water for the solution should be immersed in a vessel of cold water, since the absorption of hydrochloric acid by water is attended by the evolution of great heat; and, further, the...



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